

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Pd: \_\_\_\_\_

## Periodic Table Review Worksheet

### Chapter 3.2

1. Why do we use the periodic Table? What function does it serve?
2. What is a *Group* in the Periodic Table? What does an element's group tell you about the element?
3. What is the *atomic number*? What does it signify?
4. What is a *period* in the periodic table? What does the period tell us about an element?
5. Is an element in the 6<sup>th</sup> group more or less reactive than an element in the 1<sup>st</sup> group? Why?
6. What do all of the elements in the 18<sup>th</sup> group have in common?

7. What is an *Ion*? How does an *ion* form?
  
8. What is the name of an ion with a negative charge?
9. What is the name of an ion with a positive charge?
10. What is the charge of an atom of Carbon that has gained two electrons?
11. What is the charge of an atom of Beryllium that has lost two electrons?
12. What is an isotope?
  
13. How many neutrons does sulfur-34 have? (hint – 34 is the mass #, use the periodic table to find the atomic number).
14. How many neutrons does calcium-47 have?
15. What is the *Atomic Mass Unit*? Why do we use it?
  
16. What is the *average atomic mass*?